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# GCSE CHEMISTRY

CH3FP  
Mark scheme

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4402  
June 2014

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Version: 1.0 Final

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Mark schemes are prepared by the Lead Assessment Writer and considered, together with the relevant questions, by a panel of subject teachers. This mark scheme includes any amendments made at the standardisation events which all associates participate in and is the scheme which was used by them in this examination. The standardisation process ensures that the mark scheme covers the students' responses to questions and that every associate understands and applies it in the same correct way. As preparation for standardisation each associate analyses a number of students' scripts: alternative answers not already covered by the mark scheme are discussed and legislated for. If, after the standardisation process, associates encounter unusual answers which have not been raised they are required to refer these to the Lead Assessment Writer.

It must be stressed that a mark scheme is a working document, in many cases further developed and expanded on the basis of students' reactions to a particular paper. Assumptions about future mark schemes on the basis of one year's document should be avoided; whilst the guiding principles of assessment remain constant, details will change, depending on the content of a particular examination paper.

Further copies of this Mark Scheme are available from [aqa.org.uk](http://aqa.org.uk)

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## Information to Examiners

### 1. General

The mark scheme for each question shows:

- the marks available for each part of the question
- the total marks available for the question
- the typical answer or answers which are expected
- extra information to help the Examiner make his or her judgement and help to delineate what is acceptable or not worthy of credit or, in discursive answers, to give an overview of the area in which a mark or marks may be awarded
- the Assessment Objectives and specification content that each question is intended to cover.

The extra information is aligned to the appropriate answer in the left-hand part of the mark scheme and should only be applied to that item in the mark scheme.

At the beginning of a part of a question a reminder may be given, for example: where consequential marking needs to be considered in a calculation; or the answer may be on the diagram or at a different place on the script.

In general the right-hand side of the mark scheme is there to provide those extra details which confuse the main part of the mark scheme yet may be helpful in ensuring that marking is straightforward and consistent.

### 2. Boldening and underlining

- 2.1** In a list of acceptable answers where more than one mark is available ‘any **two** from’ is used, with the number of marks boldened. Each of the following bullet points is a potential mark.
- 2.2** A bold **and** is used to indicate that both parts of the answer are required to award the mark.
- 2.3** Alternative answers acceptable for a mark are indicated by the use of **or**. Different terms in the mark scheme are shown by a / ; eg allow smooth / free movement.
- 2.4** Any wording that is underlined is essential for the marking point to be awarded.

### 3. Marking points

#### 3.1 Marking of lists

This applies to questions requiring a set number of responses, but for which students have provided extra responses. The general principle to be followed in such a situation is that ‘right + wrong = wrong’.

Each error / contradiction negates each correct response. So, if the number of error / contradictions equals or exceeds the number of marks available for the question, no marks can be awarded.

However, responses considered to be neutral (indicated as \* in example 1) are not penalised.

Example 1: What is the pH of an acidic solution? (1 mark)

Student	Response	Marks awarded
1	green, 5	0
2	red*, 5	1
3	red*, 8	0

Example 2: Name two planets in the solar system. (2 marks)

Student	Response	Marks awarded
1	Neptune, Mars, Moon	1
2	Neptune, Sun, Mars, Moon	0

### 3.2 Use of chemical symbols / formulae

If a student writes a chemical symbol / formula instead of a required chemical name, full credit can be given if the symbol / formula is correct and if, in the context of the question, such action is appropriate.

### 3.3 Marking procedure for calculations

Full marks can be given for a correct numerical answer, without any working shown.

However, if the answer is incorrect, mark(s) can be gained by correct substitution / working and this is shown in the 'extra information' column or by each stage of a longer calculation.

### 3.4 Interpretation of 'it'

Answers using the word 'it' should be given credit only if it is clear that the 'it' refers to the correct subject.

### 3.5 Errors carried forward

Any error in the answers to a structured question should be penalised once only.

Papers should be constructed in such a way that the number of times errors can be carried forward are kept to a minimum. Allowances for errors carried forward are most likely to be restricted to calculation questions and should be shown by the abbreviation e.c.f. in the marking scheme.

### 3.6 Phonetic spelling

The phonetic spelling of correct scientific terminology should be credited **unless** there is a possible confusion with another technical term.

### 3.7 Brackets

(.....) are used to indicate information which is not essential for the mark to be awarded but is included to help the examiner identify the sense of the answer required.

### 3.8 Ignore / Insufficient / Do not allow

Ignore or insufficient is used when the information given is irrelevant to the question or not enough to gain the marking point. Any further correct amplification could gain the marking point.

Do **not** allow means that this is a wrong answer which, even if the correct answer is given, will still mean that the mark is not awarded.

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## Quality of Written Communication and levels marking

In Question 8(b)(ii) students are required to produce extended written material in English, and will be assessed on the quality of their written communication as well as the standard of the scientific response.

Students will be required to:

- use good English
- organise information clearly
- use specialist vocabulary where appropriate.

The following general criteria should be used to assign marks to a level:

### Level 1: Basic

- Knowledge of basic information
- Simple understanding
- The answer is poorly organised, with almost no specialist terms and their use demonstrating a general lack of understanding of their meaning, little or no detail
- The spelling, punctuation and grammar are very weak.

### Level 2: Clear

- Knowledge of accurate information
- Clear understanding
- The answer has some structure and organisation, use of specialist terms has been attempted but not always accurately, some detail is given
- There is reasonable accuracy in spelling, punctuation and grammar, although there may still be some errors.

### Level 3: Detailed

- Knowledge of accurate information appropriately contextualised
- Detailed understanding, supported by relevant evidence and examples
- Answer is coherent and in an organised, logical sequence, containing a wide range of appropriate or relevant specialist terms used accurately.
- The answer shows almost faultless spelling, punctuation and grammar.

Question	Answers	Extra information	Mark	AO / Spec. Ref.	ID
1(a)(i)	atomic weight		1	1 / 3.1.1a	A
1(a)(ii)	groups		1	1 / 3.1.1a	A
1(a)(iii)	left a gap		1	1 / 3.1.1b	A
1(a)(iv)	had not been discovered by 1869		1	3 / 3.1.1b	A
1(b)	protons electrons	<i>must be in correct order</i>	1 1	1 / 3.1.2a/b	G
1(c)	sodium and nickel are both metals sodium is more reactive than nickel		1 1	1 / 3.1.3c	A
1(d)(i)	bromine	allow Br <sub>2</sub> / Br do <b>not</b> allow bromide	1	2 / 3.1.3g	E
1(d)(ii)	iodine is less reactive (than bromine)	<i>it = iodine</i> allow converse do <b>not</b> allow bromide	1	1 / 3.1.3g	E
<b>Total</b>			<b>10</b>		

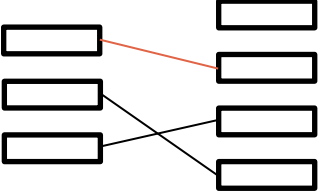
Question	Answers	Extra information	Mark	AO / Spec. Ref.	ID
2(a)	B D A (C)	allow <b>one</b> mark for one letter in correct position	2	2 / 3.2.1b/e	G
2(b)	calcium ions		1	1 / 3.2.1b	A
2(c)(i)	ethanoic acid		1	1 / 3.6.1c	A
2(c)(ii)	$  \begin{array}{ccccccc}  & & \text{H} & & \text{O} & & \\  & &   & &    & & \\  \text{H} & - & \text{C} & - & \text{C} & - & \text{O} - \text{H} \\  & &   & & & & \\  & & \text{H} & & & &   \end{array}  $		1	2 / 3.6	E
2(c)(iii)	carbon dioxide	allow CO <sub>2</sub>	1	1 / 3.4.1d	G
2(d)	more efficient heating	allow saves energy allow takes less time to boil	1	1 / 3.2.1e	E
2(e)	scum	accept calcium stearate accept magnesium stearate accept (white) precipitate	1	1 / 3.2.1a	G
<b>Total</b>			<b>8</b>		

Question	Answers	Extra information	Mark	AO / Spec. Ref.	ID
3(a)(i)	burette		1	1 / 3.4.1g	A
3(a)(ii)	indicator		1	1 / 3.4.1g	A
3(a)(iii)	colour change		1	1 / 3.4.1g	E
3(b)(i)	any <b>one</b> from: <ul style="list-style-type: none"> <li>• volume of (hydrochloric) acid</li> <li>• concentration of (hydrochloric) acid</li> <li>• concentration of (sodium) hydroxide</li> </ul>	<p><i>allow amount of (hydrochloric) acid</i></p> <p>allow concentration of alkali</p>	1	3 / 3.4.1g	E
3(b)(ii)	22.3(0)		1	2 / 3.4.1g	G
<b>Total</b>			<b>5</b>		

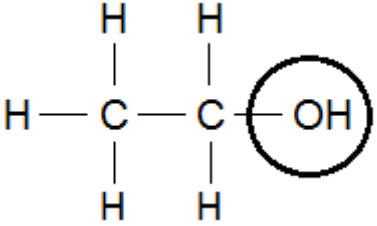


Question	Answers	Extra information	Mark	AO / Spec. Ref.	ID
4(a)	⇌ ammonia	allow NH <sub>3</sub>	1 1	2 / 3.5.1b	E
4(b)	increases quickly at first <i>then slows</i>  <i>at any number in range from 160 – 220 (atmospheres)</i>	<i>ignore levels off</i> <i>allow rate of increase slows for first two marking points</i> <i>allow any number in range 60 – 66 (%)</i>	1 1 1	2 / 3.5	E
4(c)	(nitrogen and hydrogen) recycled	<i>allow (nitrogen and hydrogen) reused</i>	1	1 / 3.5.1b	E
4(d)(i)	<i>jobs lost</i>	accept mines closed <b>or</b> local economy damaged	1	3 / 3.5	E
4(d)(ii)	any <b>one</b> from: <ul style="list-style-type: none"> <li><i>nitrates / fertilisers cost less</i></li> <li><i>more crops / food can be grown</i></li> <li><i>food costs less</i></li> <li><i>nitrates / fertilisers more widely available</i></li> </ul>		1	3 / 3.5	E
<b>Total</b>			<b>8</b>		

Question	Answers	Extra information	Mark	AO / Spec. Ref.	ID
5(a)	electrical		1	1 / 3.3.1i	A
5(b)(i) view with graph	900	accept any answer between 840 and 960	1	2 / 3.3	G
5(b)(ii)	any <b>one</b> from: <ul style="list-style-type: none"> <li>• little demand</li> <li>• few hydrogen cars</li> <li>• <i>changeover from petrol to hydrogen will take time</i></li> </ul>	<i>allow answers in terms of petrol</i>	1	3 / 3.3	E
5(c)	X on rising section of <i>line</i>		1	2 / 3.3.1d	E
<b>Total</b>			<b>4</b>		

Question	Answers	Extra information	Mark	AO / Spec. Ref.	ID
<b>6(a)(i)</b>	<i>method of introducing sample into flame</i>	<i>e.g. wire/splint/spray</i>	1	1 / 3.4	E
	<i>clean wire or colourless flame</i>	<i>allow blue / roaring flame</i>	1		
<b>6(a)(ii)</b>			1 1	2 / 3.4.1a	G
<b>6(a)(iii)</b>	(potassium) chloride	allow KCl <b>or</b> Cl <sup>-</sup>	1	2 / 3.4.1e	E
<b>6(b)(i)</b>	copper	allow Cu <sup>2+</sup>	1	2 / 3.4.1c	E
<b>6(b)(ii)</b>	sulfate		1	2 / 3.4.1f	A
<b>Total</b>			<b>7</b>		

Question	Answers	Extra information	Mark	AO / Spec. Ref.	ID
7(a)	filter to remove <u>solids</u> <b>or</b> <i>insoluble particles</i> <b>OR</b> add coagulant (1) flocculation / settling / remove solids (1)		1 1	1 / 3.2.2a/c	E
	(add) chlorine to reduce the number of microbes	accept ozone / UV accept to kill microbes / bacteria / germs accept sterilise allow disinfect ignore remove microbes	1 1		
7(b)(i)	ion exchange resin	allow ion exchange column allow sodium <u>ions</u> / Na <sup>+</sup> allow hydrogen <u>ions</u> / H <sup>+</sup>	1	1 / 3.2.2b	E
7(b)(ii)	prevent growth of microbes	accept sterilise accept to kill microbes / bacteria / germs accept to reduce the number of microbes ignore remove microbes	1	1 / 3.2.2b	E
7(c)	high cost of energy / heating	allow uses a lot of energy	1	2 / 3.2.2d	E
7(d)	any <b>one</b> from: <ul style="list-style-type: none"> <li>• helps to develop / maintain bones</li> <li>• helps to develop / maintain teeth</li> <li>• reduces heart disease</li> </ul>	allow any suitable positive effect on bones allow any suitable positive effect on teeth	1	1 / 3.2.1f	E
<b>Total</b>			<b>8</b>		

Question	Answers	Extra information	Mark	AO / Spec. Ref.	ID
8(a)(i)			1	1 / 3.6	G
8(a)(ii)	ethanol	<i>allow ethyl alcohol</i> <i>do <b>not</b> accept ethanal</i> <i>ignore all formulae</i>	1	1 / 3.3.1a	G
8(b)(i)	any <b>two</b> from: <ul style="list-style-type: none"> <li>• lid</li> <li>• <i>metal</i> calorimeter</li> <li>• insulation (around sides of beaker)</li> <li>• excluding draughts</li> <li>• stirrer</li> </ul>	<i>allow metal beaker</i> <i>do <b>not</b> allow flammable insulation / beaker</i>  <i>allow stirring</i>	2	3 / 3.3.1a	E

Question	Answers	Extra information	Mark	AO / Spec. Ref.	ID
<p><b>8(b)(ii)</b> Marks awarded for this answer will be determined by the Quality of Written Communication (QWC) as well as the standard of the scientific response. Examiners should also refer to the information on page 5, and apply a 'best-fit' approach to the marking.</p>				1 + 3 / 3.3.1a	E
<b>0 marks</b>	<b>Level 1 (1–2 marks)</b>	<b>Level 2 (3–4 marks)</b>	<b>Level 3 (5–6 marks)</b>		
No relevant content	<i>There is a description of part of an experimental method <b>or</b> a measurement which should be taken.</i>	<i>There is a description of some parts of an experimental method <b>and</b> a measurement which should be taken.</i>	<i>There is a description of an experimental method <b>and</b> measurements which should be taken.</i>		
<b>Examples of the point that may be made in the response</b>					
<ul style="list-style-type: none"> <li>• <i>light ethanol and heat water</i></li> <li>• <i>extinguish ethanol</i></li> <li>• <i>after suitable temperature rise <b>or</b> after a suitable time</i></li> <li>• <i>stir water</i></li> <li>• <i>measure mass / volume of water</i></li> <li>• <i>measure initial temperature of water</i></li> <li>• <i>measure final temperature of water</i></li> <li>• <i>measure temperature rise</i></li> <li>• <i>measure initial mass of ethanol (and burner)</i></li> <li>• <i>measure final mass of ethanol (and burner)</i></li> <li>• <i>measure change in mass of ethanol</i></li> </ul>					
<b>Total</b>			<b>9</b>		